

What's Going On?

Checking In

Minds on

Homework Groups

Action!

Molecular Compounds

Consolidation

BBC Interactive Lesson

**Learning Goal - I will be able to describe how
molecular compounds are formed.**

 Minds on

Homework Groups

Pg. 221 - 225

All questions

Pg. 221

Write the name of the compound

1. NaF - Sodium fluoride

2. KI - Potassium iodide

3. MgCl₂ - Magnesium chloride4. AlCl₃ - Aluminum chloride5. Ca₃P₂ - Calcium phosphide

Pg. 222

Write the name of the compound

1. FeCl_3 Iron (III) chloride

2. PbO_2 Lead (IV) oxide

3. Ni_2S_3 Nickel (III) sulfide

4. CuF_2 Copper (II) fluoride

5. Cr_2S_3 Chromium (III) sulfide

Pg. 223

Write the name of the compound

1. KOH

Potassium hydroxide

2. ZnCO₃

Zinc carbonate

3. Mg₃(PO₄)₂

Magnesium phosphate

4. CaSO₄

Calcium sulfate

5. Al₂(CO₃)₃

Aluminum carbonate

Pg. 224

Write the formula

1. lithium bromide LiBr

2. magnesium fluoride



3. silver nitride

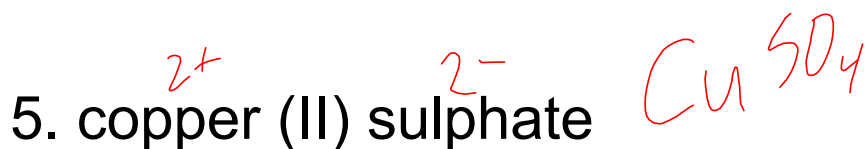
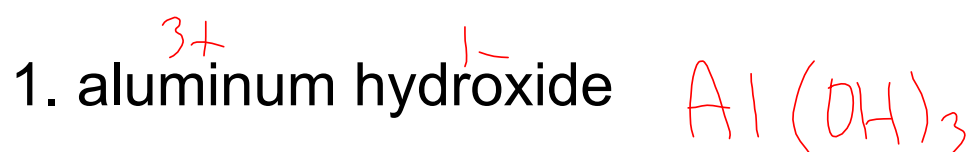
4. iron (III) chloride FeCl_3

5. chromium (III) sulphide



Pg. 225

Write the formula



Action!

Molecular Compounds

When non-metals combine, a pure substance called a molecular compound is formed.

In molecular compounds, atoms share electrons to form small groups, called molecules.

Action!

Molecular Compounds

Most molecular compounds share the following properties:

- can be solids, liquid or gas at room temperature
- usually good insulators but poor conductors of electricity
- have relatively low boiling points

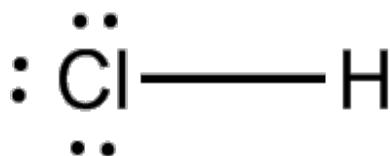
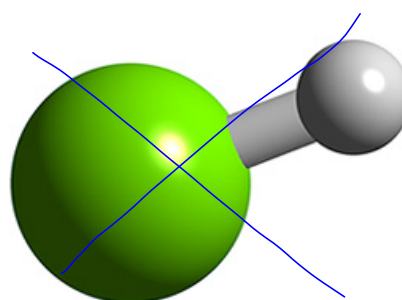
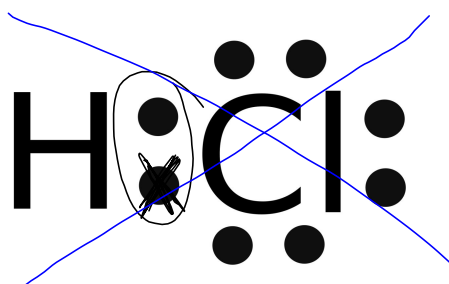
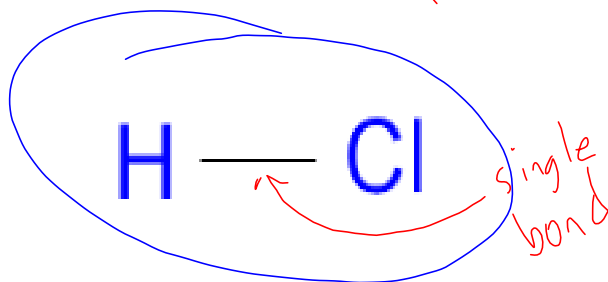
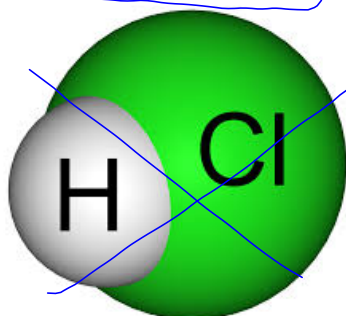
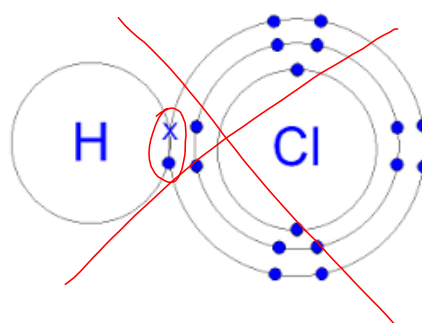
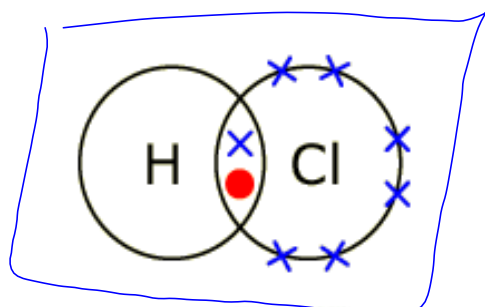
Action!

Molecular Compounds

Examples

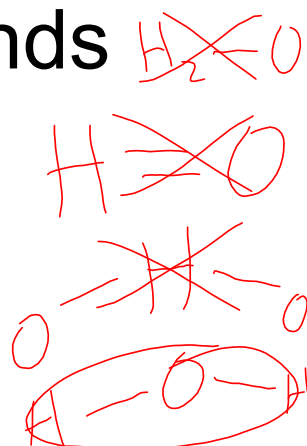
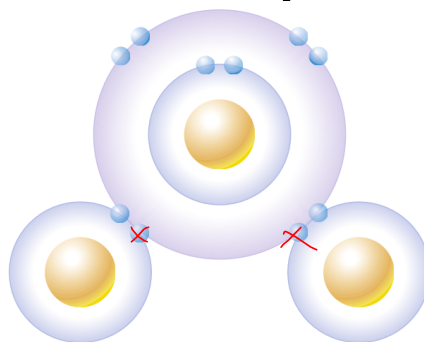
Hydrogen chloride (HCl)

Hydrochloric acid



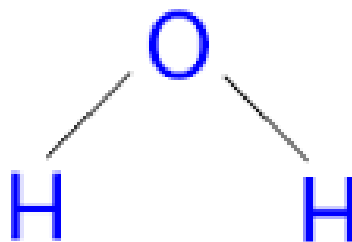
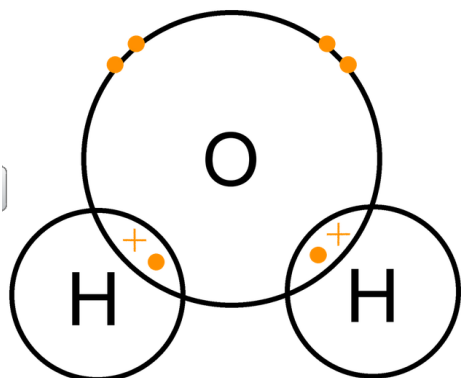
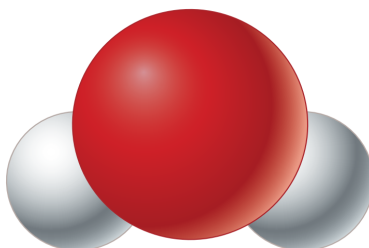
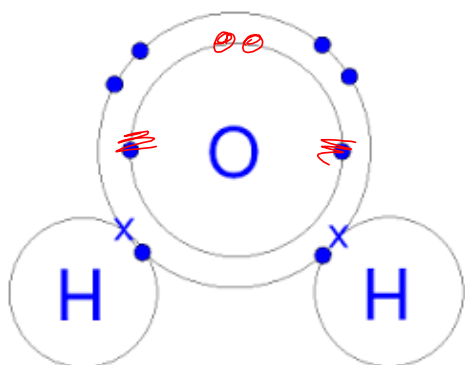
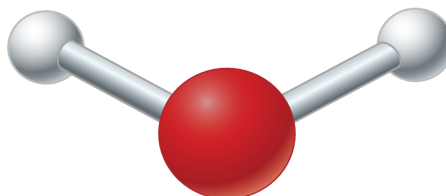
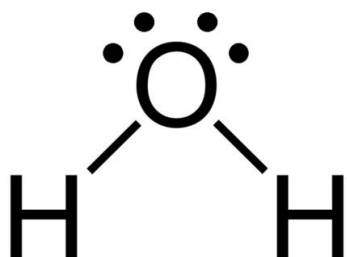
Action!

Molecular Compounds



Examples

Water (H_2O)



Action!

Molecular Compounds

Naming Molecular Compounds

1. Name the first element.
 2. Name the second element using the ending "-ide"
 3. Add prefixes to each element to indicate the number of atoms of each element.
- *We do not add "mono" to the first element in a compound.

Prefixes for Molecular Compound Naming	
Number of Atoms	Prefix
1	mono
2	di
3	tri
4	tetra
5	penta
6	hexa

Consolidation

BBC Interactive Lesson

Consolidation

Practice

Pg. 217

1 - 10

Pg. 229

1 - 9

Attachments



1D CHEM - B1 (History of the Atom) - Atomic Theory.mp4